

CHEMISTRY

Paper 1 Multiple Choice

5070/12 October/November 2010 1 hour

Additional Materials:

Multiple Choice Answer Sheet Soft clean eraser Soft pencil (type B or HB recommended)

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

04

Do not use staples, paper clips, highlighters, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A**, **B**, **C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer. Any rough working should be done in this booklet. A copy of the Periodic Table is printed on page 20.

This document consists of 18 printed pages and 2 blank pages.



1 The boiling points of various gases found in the air are shown below.

	°C
argon	-186
carbon dioxide	-78
nitrogen	-198
oxygen	-183

If the air is cooled, the first substance to condense is water.

If the temperature is lowered further, what is the next substance to condense?

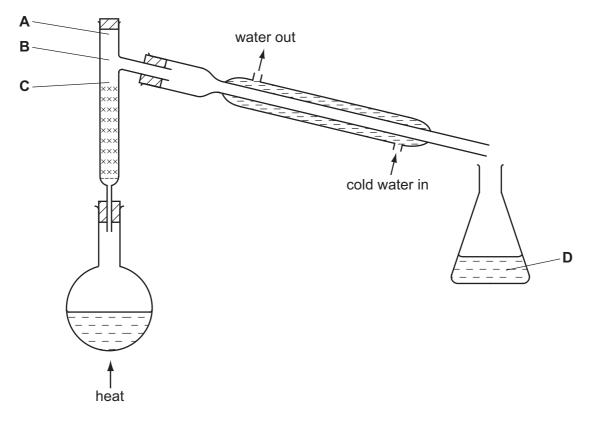
- A argon
- B carbon dioxide
- **C** nitrogen
- D oxygen
- 2 Substance X dissolves in water to form a colourless solution. This solution reacts with aqueous lead(II) nitrate in the presence of dilute nitric acid to give a yellow precipitate.

What is substance X?

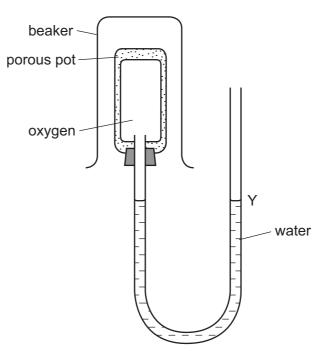
- A calcium iodide
- **B** copper(II) chloride
- **C** iron(II) iodide
- D sodium chloride

3 The fractional distillation apparatus shown is to be used for separating a mixture of two colourless liquids. A thermometer is missing from the apparatus.

Where should the bulb of the thermometer be placed?



4 The diagram shows a diffusion experiment.



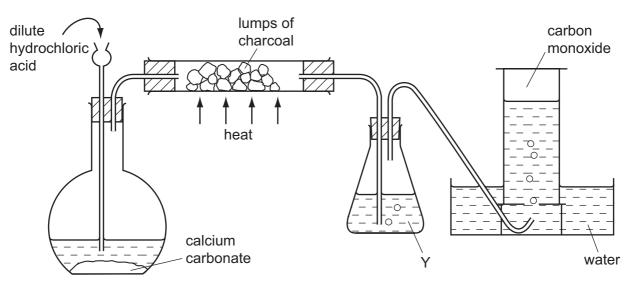
Which gas, when present in the beaker over the porous pot, will cause the water level at Y to rise?

- A carbon dioxide, CO₂
- **B** chlorine, Cl_2
- C methane, CH₄
- D nitrogen dioxide, NO₂
- **5** Hydrogen can form both H^+ ions and H^- ions.

Which one of the statements below is correct?

- **A** An H^+ ion has more protons than an H^- ion.
- **B** An H^+ ion has no electrons.
- **C** An H^- ion has one more electron than an H^+ ion.
- **D** An H^- ion is formed when a hydrogen atom loses an electron.

6 The diagram shows apparatus used to obtain carbon monoxide.



What is the main purpose of Y?

- **A** to dry the gas
- **B** to prevent water being sucked back on to the hot carbon
- **C** to remove carbon dioxide from the gas
- **D** to remove hydrogen chloride from the gas
- 7 A dark, shiny solid, X, conducts electricity.

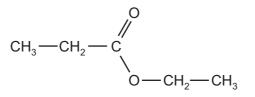
Oxygen combines with X to form a gaseous oxide.

What is X?

- A graphite
- B iodine
- **C** iron
- D lead
- 8 Which substance could be sodium chloride?

	molting point /°C	conduction	of electricity
	melting point/°C	when liquid	in aqueous solution
Α	-114	nil	good
В	180	nil	nil (insoluble)
С	808	good	good
D	3550	nil	nil (insoluble)

9 The diagram shows the molecule ethyl propanoate.

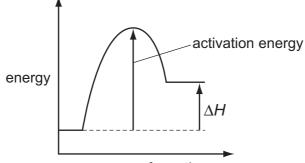


How many bonding pairs of electrons are there in the molecule?

A 13 **B** 16 **C** 17 **D** 20

- 10 The conduction of electricity by metals is carried out by the movement of
 - A electrons only.
 - **B** electrons and positive ions.
 - C negative ions only.
 - **D** negative ions and positive ions.
- **11** What is the concentration of iodine molecules, I_2 , in a solution containing 2.54 g of iodine in 250 cm³ of solution?
 - **A** 0.01 mol/dm³
 - **B** $0.02 \text{ mol}/\text{dm}^3$
 - C 0.04 mol/dm³
 - **D** $0.08 \text{ mol}/\text{dm}^3$

12 The energy profile for the forward direction of a **reversible** reaction is shown.

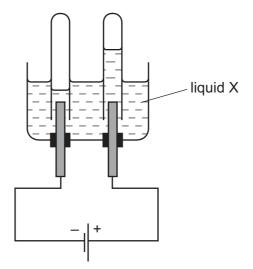


progress of reaction

Which row correctly shows the sign of both the activation energy and the type of the enthalpy change for the **reverse** reaction?

	sign of activation energy	type of enthalpy change
Α	negative	endothermic
В	negative	exothermic
С	positive	endothermic
D	positive	exothermic

13 The diagram shows the results of an electrolysis experiment using inert electrodes.



Which could be liquid X?

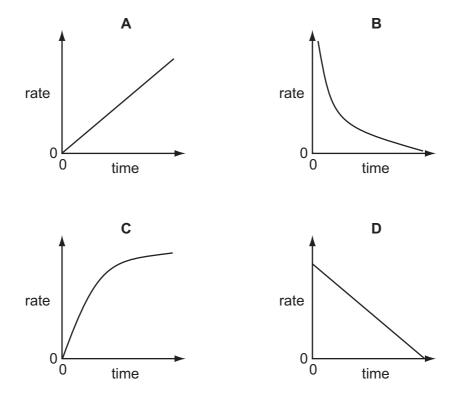
- A aqueous copper(II) sulfate
- B concentrated aqueous sodium chloride
- C dilute sulfuric acid
- D ethanol

- 14 In which reaction is nitric acid acting as an oxidising agent?
 - $\textbf{A} \quad Cu + 4HNO_3 \rightarrow Cu(NO_3)_2 + 2H_2O + 2NO_2$
 - **B** CuO + 2HNO₃ \rightarrow Cu(NO₃)₂ + H₂O
 - $\textbf{C} \quad Na_2CO_3 + 2HNO_3 \rightarrow 2NaNO_3 + H_2O + CO_2$
 - **D** NaOH + HNO₃ \rightarrow NaNO₃ + H₂O
- 15 The equation shows the formation of sulfur trioxide in the Contact process.

 $2SO_2(g) + O_2(g) \rightleftharpoons 2SO_3(g) \quad \Delta H = -95 \text{ kJ/mol}$

What would decrease the yield of sulfur trioxide in a given time?

- A addition of more oxygen
- B an increase in pressure
- C an increase in temperature
- **D** removal of $SO_3(g)$ from the reaction chamber
- **16** Which graph represents how the rate of reaction varies with time when an excess of calcium carbonate reacts with dilute hydrochloric acid?



17 The tests below were carried out on a solution containing ions of the metal X.

test	observation
add sodium chloride solution	no change
add sodium sulfate solution	no change
add sodium hydroxide solution	a precipitate was formed, soluble in excess of the hydroxide

What is metal X?

- A calcium
- B iron
- C lead
- D zinc
- **18** A student mixed together aqueous solutions of Y and Z. A white precipitate formed.

Which could **not** be solutions Y and Z?

	solution Y	solution Z
Α	hydrochloric acid	silver nitrate
В	hydrochloric acid	sodium nitrate
С	sodium chloride	lead(II) nitrate
D	sodium chloride	silver nitrate

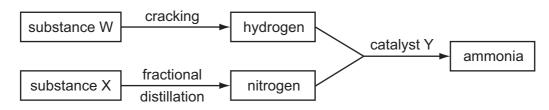
19 Sulfur is burnt in air.

Which statement about this reaction is correct?

- **A** Sulfur is oxidised to sulfur trioxide.
- **B** The gas formed turns aqueous potassium dichromate(VI) from orange to green.
- **C** The reaction is reversible.
- **D** The reaction needs a catalyst.
- 20 Which property is common to calcium, potassium and sodium?
 - A Their atoms all lose two electrons when they form ions.
 - **B** They all form carbonates which are insoluble in water.
 - **C** They are all less dense than water.
 - **D** They are all metallic.

- 21 Which set of the electronic structures are only found in metals?
 - 2, 1 2, 8, 1 2, 8, 8, 1 Α 2, 5 2, 6 2, 7 В 2, 7 2, 8, 7 С 2, 8, 18, 7 D 2, 8, 3 2, 8, 4 2, 8, 5

22 The diagram shows processes that take place in the manufacture of ammonia.



What are substances W and X and catalyst Y?

	W	Х	Y
Α	air	oil	iron
в	air	oil	vanadium(V) oxide
С	oil	air	iron
D	oil	air	vanadium(V) oxide

23 The position of metal M in the reactivity series is shown.

decrease in reactivity

K, Na, M, Al, Zn, Fe, Pb, Cu, Ag

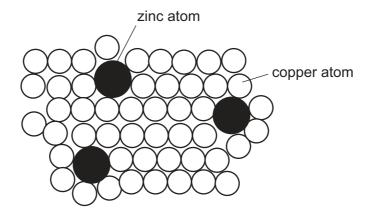
Which method will be used to extract M from its ore?

- A electrolysis of its aqueous sulfate
- B electrolysis of its molten oxide
- C reduction of its oxide by heating with coke
- D reduction of its oxide by heating with hydrogen

24 When zinc is added to a solution of a metal sulfate, the metal is deposited and zinc ions are produced in solution.

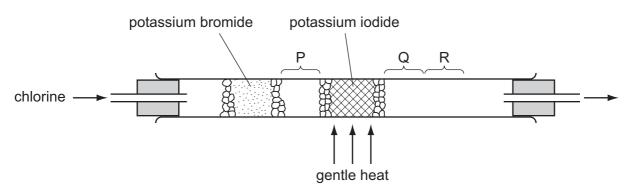
Which metal is deposited?

- A calcium
- B copper
- **C** magnesium
- D potassium
- 25 The diagram shows the structure of brass.



Why is brass harder than pure copper?

- A The zinc atoms form strong covalent bonds with copper atoms.
- **B** The zinc atoms prevent layers of copper atoms from slipping over each other easily.
- **C** The zinc atoms prevent the 'sea of electrons' from moving freely in the solid.
- **D** Zinc atoms have more electrons than copper atoms.



26 Using the apparatus shown, chlorine is passed through the tube.

After a short time, coloured substances are seen at P, Q and R.

What are these coloured substances?

	at P	at Q	at R
Α	green gas	red brown vapour	violet vapour
в	green gas	violet vapour	black solid
С	red brown vapour	violet vapour	black solid
D	violet vapour	red brown vapour	red brown vapour

- **27** In the electrolysis of molten aluminium oxide for the extraction of aluminium, the following three reactions take place.
 - $1 \quad \mathsf{A}l^{3^+} + 3e^- \to \mathsf{A}l$

$$2 \quad 20^{2-} \rightarrow O_2 + 4e^-$$

 $3 \quad C + O_2 \rightarrow CO_2$

Which reactions take place at the anode?

Α	1 only	в	2 only	С	1 and 3	D	2 and 3
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28 Which equation in the blast furnace extraction of iron is not a redox reaction?

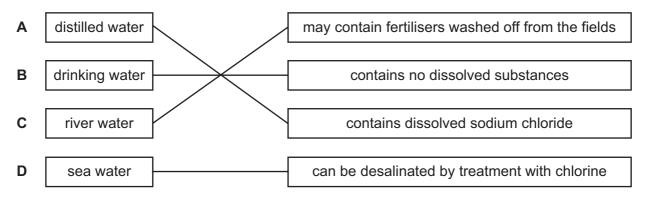
A
$$CaCO_3 \rightarrow CaO + CO_2$$

- $\textbf{B} \quad 2C + O_2 \rightarrow 2CO$
- $\textbf{C} \quad C + CO_2 \rightarrow 2CO$
- $\textbf{D} \quad \text{Fe}_2\text{O}_3 \textbf{+} 3\text{CO} \rightarrow 2\text{Fe} \textbf{+} 3\text{CO}_2$

29 Which statement about the material used for aircraft bodies is correct?

Aircraft bodies are made from

- A an aluminium alloy because pure aluminium is too soft.
- **B** pure aluminium because of its high melting point.
- **C** pure aluminium because of its low density.
- **D** pure aluminium because of its resistance to corrosion.
- 30 Which natural process can cause nitrogen oxides to be formed in the atmosphere?
 - A bacterial decay of plants
 - B lightning activity
 - **C** photosynthesis
 - **D** respiration
- **31** Which type of water in the left hand column is linked correctly to a statement in the right hand column?

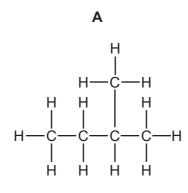


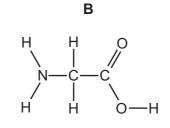
32 A catalytic converter in a car exhaust system speeds up the change of pollutants into less harmful products.

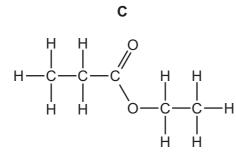
Which change does not occur in a catalytic converter?

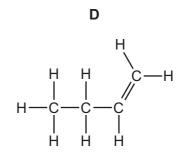
- **A** carbon dioxide \rightarrow carbon
- $\textbf{B} \quad \text{carbon monoxide} \rightarrow \text{carbon dioxide}$
- **C** nitrogen oxides \rightarrow nitrogen
- $\textbf{D} \quad \text{unburned hydrocarbons} \rightarrow \text{carbon dioxide and water}$

33 Which formula represents a compound likely to undergo addition polymerisation?





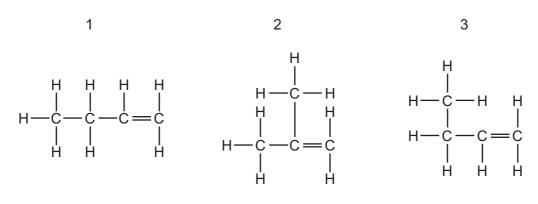


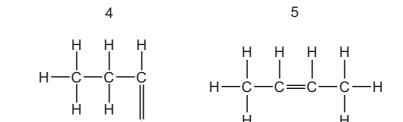


- 34 Which statement about ethanol is correct?
 - A It is an unsaturated compound.
 - **B** It is formed by the catalytic addition of steam to ethene.
 - **C** It is formed by the oxidation of ethanoic acid.
 - **D** It reacts with ethyl ethanoate to form an acid.
- **35** An organic compound has an empirical formula C₂H₄O.

What is the compound?

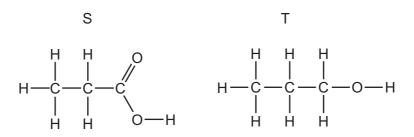
- A butanoic acid
- B butanol
- **C** ethanoic acid
- D ethanol





Which structures represent identical molecules?

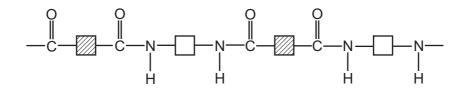
- A 1 and 3 only
- B 2 and 3 only
- **C** 1, 3 and 4 only
- **D** 1, 3 and 5 only
- 37 The diagrams show two organic compounds.



Which statement about the compounds S and T is correct?

- **A** Both S and T react with sodium carbonate.
- **B** S and T react together to form the ester ethyl propanoate.
- **C** T can be changed into S using acidified potassium dichromate(VI).
- **D** They are in the same homologous series.

38 Polymer X has the structure shown.



The list shows four terms that can be applied to polymers.

- 1 addition polymer
- 2 condensation polymer
- 3 polyamide
- 4 polyester

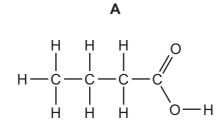
Which two terms can be applied to polymer X?

A 1 and 3 **B** 1 and 4 **C** 2 and 3 **D** 2 and 4

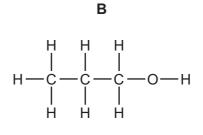
- 39 In which reaction is water produced?
 - A manufacture of ethanol from ethene
 - B manufacture of margarine from vegetable oils
 - **C** manufacture of poly(ethene) from ethene
 - **D** manufacture of *Terylene* from a carboxylic acid and an alcohol

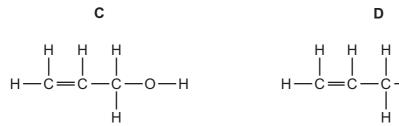
test result add bromine water turns colourless add aqueous sodium carbonate carbon dioxide formed

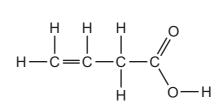
What is compound Z?



40 The results of tests on compound Z are shown.







D

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19

DATA SHEET The Periodic Table of the Elements

							5	Group								
											≡	≥	>	N	١١	0
						-										4
						т										He
						Hydrogen 1										Aelium 2
	6										1	12	14	16	19	20
	Be										ш	ပ	z	0	ш	Ne
Lithium Ber	Beryllium										5 5	Carbon 6	Nitrogen 7	Oxygen 8	Fluorine 9	Neon 10
23	24										27	28	31	32	35.5	40
	Mg										٩l	Si	٩	S	C1	Ar
Sodium Magr 12	Magnesium 12										Aluminium 13	Silicon 14	Phosphorus 15		Chlorine 17	Argon 18
	40 4	45 48	51	52	55	56	59	59	64	65	70	73	75	5	80	84
×	Ca	Sc Ti	>	ں م	Mn	Fe	ပိ	Ni	Cu	Zn	Ga	e	As	Se	Ŗ	Кr
Potassium Cal 19 20	0	Scandium Titanium	Vanadium 23	Chromium 24	Manganese 25	lron 26	Cobalt 27	Nickel 28	Copper 29	Zinc 30	Gallium 31	E	Arsenic 33	Selenium 34	Bromine 35	Krypton 36
85			93	96		101	103	106	108	112	115		122	128	127	131
Rb		Y Zr		Mo	Тc	Ru	Rh	Pd	Ag	Cd	In		Sb	Te	I	Xe
Rubidium Strontium 38	36	Yttrium Zirconium	Niobium 41	Molybdenum 42	Technetium 43	Ruthenium 44	Rhodium 45	Palladium 46	Silver 47	Cadmium 48	Indium 49	Tin 50	Antimony 51	Tellurium 52	lodine 53	Xenon 54
		139 178		184	186	190	192	195	197	201	204		209			
			Та	>	Re	Os	Ir	Pt	Au	Hg	LΙ	Pb	Bi		At	Rn
56	2 2	T * T2	Tantalum 73	Tungsten 74	Rhenium 75	Osmium 76	Iridium 77	Platinum 78	Gold 79	Mercury 80	Thallium 81	Lead 82	Bismuth 83	_	Astatine 85	Radon 86
		227														
		Ac														
Francium Ra	Radium Acti 89	Actinium														
71 Lanth	*58-71 Lanthanoid series	GS	140	141	144		150	152	157	159	162	165	167	169	173	175
103 Actin	190-103 Actinoid series		Cerium Cerium	Pr Praseodymium	Neodymium 60	Promethium 61	Samarium Samarium	Eu Europium	Gd Gadolinium	Tb Terbium	Dysprosium Bysprosium	Holmium E7	Erbium Ge	Thulium	Yb Ytterbium	Lutetium 71
a	a = relati	a = relative atomic mass	232	60			70	8	5	8	8	10	8	60	2	-
Kev X	X = atom	X = atomic symbol	H H	Ра	n	aN	Pu	Am	Cm	Bk	Ç	Es	Em	Md	No	L
q	b = proto	b = proton (atomic) number	06	Protactinium 91	ε	Neptunium 93	Plutonium 94	Americium 95	Curium 96	Berkelium 97	Californium 98	Einsteinium 99	Fermium 100	Mendelevium 101	Nobelium 102	Lawrencium 103

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